

Thermo. Titr. Application Note No. H-081

Title: Determination of Phosphoric and Nitric Acid in Nitrophos Liquors

Scope: Determination of phosphoric and nitric acids in liquors from the Nitrophos fertilizer manufacturing process.

Principle: Titration in a standardized solution of 2 mol/L NaOH in a sodium chloride solution to sharpen the third endpoint.

Reagents: *Titrant:* 2 mol/L NaOH, standardized against potassium hydrogen phthalate
Conditioning solution: 180g/L NaCl

Method: Basic Experimental Parameters:

Titration delivery rate (mL/min.)	5
Delay before titration commences (sec)	5
No. of exothermic endpoints	3
Data smoothing factor (DSF)	93
Stirring speed (802 stirrer)	15

Titration: Weigh accurately approximately 0.7mL (approximately 1g) of Nitrophos liquor into a clean, dry titration vessel. Add 30mL 180g/L NaCl solution, and allow to stand for approximately 30 minutes before titrating in order that any contained hexafluorosilicic acid is fully hydrolyzed.

Blank determinations: Prepare titration solutions as above, weighing masses of Nitrophos liquor encompassing a range from approximately 1.5 to 0.6g. Titrate, and prepare two graphs. In the first graph, plot sample mass (x-axis) against titration volume to the first endpoint, EP1. Determine the y-intercept from regression analysis. This will be the blank (B1) for the nitric acid determination. In the second graph, plot sample mass (x-axis) against the difference between the third and second endpoints, EP3-EP2. Once again, determine the y-intercept from regression analysis. This will be the blank (B2) for the phosphoric acid determination

<i>Examples:</i>		<i>Nitrophos process liquor</i>	
	H ₃ PO ₄ % w/w		24.1±0.25 (n=11)
	HNO ₃ % w/w		7.8±0.17 (n=11)

Calculations:

$$\text{H}_3\text{PO}_4 \% = \frac{((\text{EP3} - \text{EP2} - \text{B2}) \times \text{NaOH mol/L} \times 97.9769 \times 100)}{(\text{sample mass, g} \times 1000)}$$

$$\text{HNO}_3 \% = \frac{(((\text{EP1} - \text{B1}) - (\text{EP3} - \text{EP2} - \text{B2})) \times \text{NaOH mol/L} \times 62.99564 \times 100)}{(\text{sample mass, g} \times 1000)}$$

Thermometric Titration Plot:

Legend:
 Red = solution temperature curve
 Black = second derivative curve (for endpoints)

